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Chapter 8 Covalent Bonding Pt 1

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Introduction : Chemical Bonding |

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Chapter 8 Covalent Bonding

242 Chapter 8 • Covalent Bonding

Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the

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Bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Pearson Education configuration. In
covalent bonding, atoms share
electrons to achieve noble gas
configuration. Most atoms share
electrons until they have a total of 8
valence electrons (octet rule). Chapter
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Chemistry Chapter 8: Covalent Bonding. Covalent Bond. Molecule. Diatomic molecule. Molecular compound. A bond formed by the sharing of electrons between atoms. A neutral group of atoms joined together by covalent bonds. A molecule consisting of two atoms.

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Bonding PDF Answers

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Combinations of atoms of the
nonmetals and metalloids in Groups
4A, 5A, 6A, and 7A of the periodic

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Bonding Answers
table are likely to form covalent bonds. • The combined atoms usually acquire a total of eight electrons, or an

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Reading every day is gratifying for many people. However, there are yet many people who along with don't similar to reading. This is a problem. But, bearing in mind you can withhold others to start reading, it will be better.

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Bonding Chapter 8 Covalent Bonding Answers

Polar Covalent Bonds • Though atoms often form compounds by sharing electrons, the electrons are not always shared equally. • Fluorine pulls harder on the electrons it shares with hydrogen than hydrogen does.

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- Therefore, the fluorine end of the molecule has more electron density than the hydrogen end.

Chapter 8 Concepts of Chemical
Bonding - Central Lyon

Chapter 8 Covalent Bonding 183

Section Review Objectives • State a

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rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and

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