

## Chemistry B Moles Packet Answer Key

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CHEMISTRY B- MOLES PACKET NAME: HR: \_\_\_\_\_ PAGE 1  
CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

Chemistry B Moles Packet - Teacher Sep  
A mole of carbon contains 6.02 x 10<sup>23</sup> atoms of carbon. A mole of helium contains 6.02 x 10<sup>23</sup> atoms of helium. A mole of sodium contains 6.02 x 10<sup>23</sup> atoms of sodium. A mole of gold contains 6.02 x 10<sup>23</sup> atoms of gold. However, we have learned about seven elements that exist as diatomic molecules:H 2, N 2, O 2, F 2, Cl 2, Br 2, and I 2

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CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 4CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASSNow that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass ofseveral moles or the mass of a fraction of a mole using the factor-label technique.1 mol = a molar mass of an atom/molecule (g/mol)EXAMPLE: What is the mass of 5.00 moles of water H 2 O?STEP1: 2 H = 2 x ( 1.01) = 2.02 STEP 2: ? grams H 2 O = 5.00 moles H 2 O x 18.02 g H 2 O = 90.10 ...

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CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 6CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBERNow that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you cancombine these two definitions into one problem.EXAMPLE: How many molecules are there in 90.1 grams of water?2 H = 2 x(1.01) = 2.02O = 1 x (16.00) = 16.00H 2 O= 18.02 g? molecules H 2 O = 90.1 g H 2 O x 1 mole H 2 O x 6.02 x 10 23 molecules H 2 O = 3.01 x 10 24 molecules H 2 ...

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